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11. molar mass = mass/moles; mass = 45.0 g; find moles;  $\Delta T_b = 90.2^\circ\text{C} - 80.1^\circ\text{C} = 10.1^\circ\text{C}$   $\left(\text{moles} = \frac{\Delta T_b}{K_f} \times \frac{\text{kg benzene}}{10.1^\circ\text{C}} \times 2.53\right)$   $\times 0.0750 = 0.300 \text{ moles}$

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11.1. Classify the following as primary, secondary and tertiary alcohols. Ans: Primary alcohols: (i), (ii), (iii) Secondary alcohols: (iv), (v) Tertiary alcohols: (vi) 11.2. Identify allylic alcohols in the above examples. Ans: (ii) and (iv) i.e. H 2 C=CH – CH 2 OH and. 11.3.

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Name SECTION 3 Date CHAPTER 11 REVIEW Gases Class SHORT ANSWER Answer the following questions in the space provided. c c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (b) divided by the mass of 1 mol. nRT (c) multiplied by 22.4 L. (d) divided by 22.4 L.

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Answer: Yes, you can find important questions for CBSE Class 11 Chemistry Chapter 2: Structure of Atom online on Vedantu. These important questions are curated and prepared by our subject matter experts, with close reference to the NCERT Chemistry textbook for Class 11 and various other reference study materials.